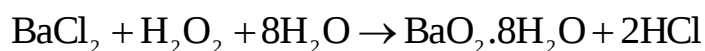


Synthesis of barium peroxide – Protocol B



Reaction. Under a nitrogen atmosphere, add a solution of hydrogen peroxide (stoichiometric proportions) and a solution of ammonium hydroxide to an aqueous solution of barium chloride. Place the mixture in an ice bath to precipitate the product.

Isolation. Collect by filtration the barium peroxide precipitated and dry at 150 °C.

Purification. Not prescribed.

Safety. See hazards associated with the reagents in Table 1.

Greenness Assessment. The evaluation was performed using the Green Star (GS) and the results are shown in Figure 1.

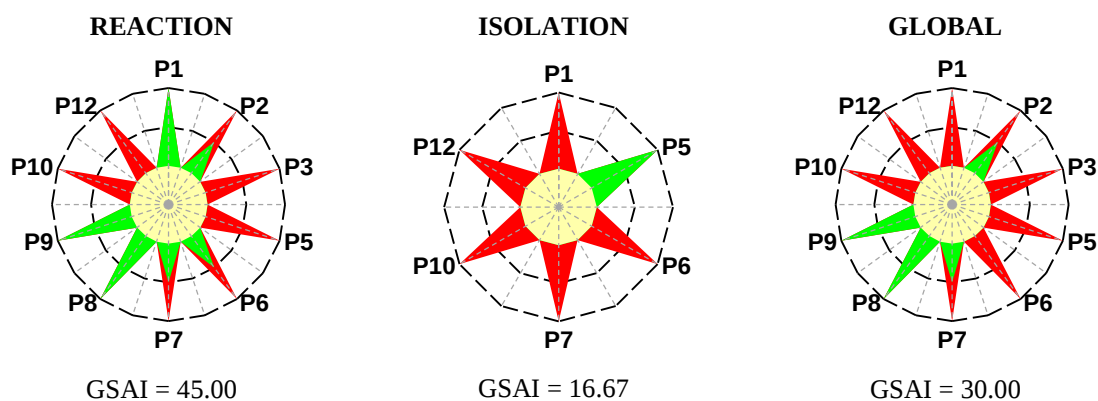


Figure 1. Greenness assessment (GS) for the synthesis of barium peroxide

Construction of the GS

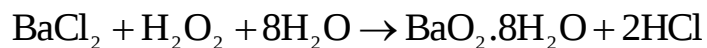


Table 1 presents the hazards and scores associated with the substances involved and Table 2 presents the scores used to construct the green stars.

Table 1. Hazards for the synthesis of barium peroxide, protocol B^α

Substances involved	Step			Hazard code	Score: hazards to...		
	R	I	Pu		HH	E	P
Reagentes estequiométricos							
Barium chloride (CAS 10361-37-2)	✓			H301, H332	3	1	1
Hydrogen peroxide (CAS 7722-84-1)	✓			H302, H318	3	1	1
Auxiliary substances							
Solvents							
Ammonium hydroxide (solution)	✓			H290, H314, H335, H400	3	3	2
Water ^{a,b}	✓			-	1	1	1
Other auxiliary substances							
Nitrogen ^a (CAS 7727-37-9)	✓			H280	1	1	2
Product							
Barium peroxide (1304-29-6)	✓	✓		H272 (cat. 2), H302, H332	2	1	3
Waste							
Ammonium hydroxide (dilute solution)		✓		H315, H318, H400	3	3	1
Hydrochloric acid (dilute solution)		✓		-	1	1	1
Nitrogen ^a	✓			H280	1	1	2
Water ^{a,b}		✓		-	1	1	1

^α R – Reaction; I – Isolation; Pu – Purification; HH – Human Health; E – Environment; P – Physical

^a Renewable; ^b Degradable to innocuous products

Table 2. Scores used to construct the green star for the synthesis of barium peroxide, protocol B^a

Green Chemistry Principle	Reaction		Isolation		Global	
	s	Explanation	s	Explanation	s	Explanation
P1 Prevention	3	Nitrogen	1	Dilute ammonium hydroxide, H318 and H400	1	Dilute ammonium hydroxide, H318 and H400
P2 Atom Economy	2	Stoichiometric proportions of reagents, formation of by-products		NA	2	Stoichiometric proportions of reagents, formation of by-products
P3 Less hazardous chemical synthesis	1	Barium chloride, H301, hydrogen peroxide, H318, and ammonium hydroxide, H314 and H400		NA	1	Barium chloride, H301, hydrogen peroxide, H318, and ammonium hydroxide, H314 and H400
P5 Safer solvents and auxiliary substances	1	Ammonium hydroxide, H314 and H400	3	Solvents and auxiliary substances are not used	1	Ammonium hydroxide, H314 and H400
P6 Increase energy efficiency	2	$0\text{ }^{\circ}\text{C} \leq T \leq 100\text{ }^{\circ}\text{C}$	1	$T > 100\text{ }^{\circ}\text{C}$	1	$T > 100\text{ }^{\circ}\text{C}$
P7 Use renewable feedstocks	2	Nitrogen is renewable	1	Substances not renewable	2	Nitrogen is renewable
P8 Reduce derivatives	3	One stage		NA	3	One stage
P9 Catalysts	3	Without catalysts		NA	3	Without catalysts
P10 Design for degradation	1	Substances not degradable	1	Substances not degradable	1	Substances not degradable
P12 Safer chemistry for accident prevention	1	Barium chloride, H301, hydrogen peroxide, H318, ammonium hydroxide, H314, and barium peroxide, H272 (cat.2)	1	Barium peroxide, H272 (cat.2), and ammonium hydroxide, H314	1	Barium chloride, H301, hydrogen peroxide, H318, ammonium hydroxide, H314, and barium peroxide, H272 (cat.2)

^as – Score; NA – Not applicable

References

Kopnin, E.M. *et al.* New Family of Au-Based Superconductors $\text{AuBa}_2\text{Ca}_{n-1}\text{Cu}_n\text{O}_{2n+3}$ ($n = 3,4$). *Chem. Mater.*, **2001**, *13*, 2905-2908.